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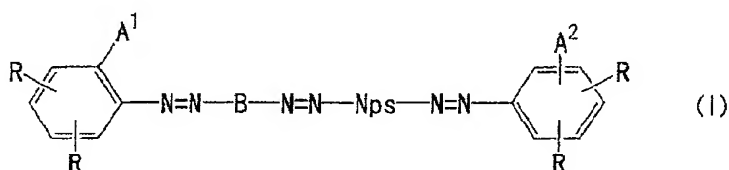
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(54) **Novel trisazo compound**

(57) The objective of the present invention is to provide a high-purity trisazo compound with excellent light resistance that is useful as a coloring matter for a black aqueous ink for ink jet recording and the like. Thus, the present invention provides a trisazo compound represented by the formula:



wherein A¹ is a group represented by the formula -COOM or -SO₃M, A³ is a hydrogen atom or a group represented by the formula -COOM or -SO₃M, M is each independently a hydrogen atom, or onium, R is each independently a hydrogen atom, an alkyl group, an alkoxy group and the like, B is a residue of K acid, Nps is a residue of γ acid, J acid, or RR acid.

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Description

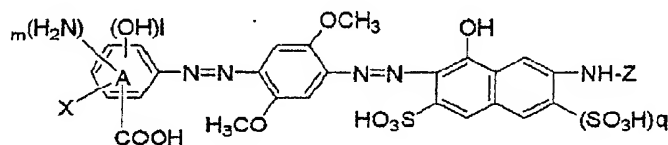
[0001] The present invention relates to a trisazo compound useful as a coloring matter for an aqueous ink for writing materials or for an aqueous ink for ink jet recording.

[0002] In the field of ink jet printing, there is a constant demand for highspeed and highly accurate printing. The level of the demand are becoming higher and higher. For example, high fastness such as light resistance and water resistance of printed materials, long-term stability of ink performance, safety and high-quality printing on various types of paper typified by plane paper and the like have recently been required. While the qualities that printed materials are required are becoming higher, the conventional dyes used as coloring matters of inks are unsatisfactory in density of printed materials and fastness with respect to water resistance and light resistance thereof, and are poor in long-term stability of ink. Thus, the conventional dyes have been becoming very difficult to simultaneously satisfy the many required qualities.

[0003] Under these circumstances, there are many attempts to satisfy the demands for density, water resistance and light resistance of printed materials by using pigments as coloring matters of printing inks. However, pigments have still not reached the levels of the dyes in print qualities and have problems with respect to long-term stability with time as ink. Therefore, pigments have not satisfied the demands sufficiently.

[0004] Further, pigments result in high cost because they require highly developed manufacturing techniques for the improvement in their physical and chemical properties such as dispersibility, stability with time and print density; pigments have disadvantage for providing low-cost products. Therefore, in the field of ink for ink jet, there is a particularly strong demand for dyes excellent in fastness with respect to density, water resistance and light resistance of printed materials.

[0005] In Japanese Patent Laid-Open Publication No. 2000-309721, a dye shown below is proposed that yields a recorded image of high density and of high toughness when printing is conducted on plain paper and that provides an ink stable for a long term.



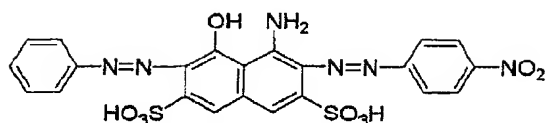
wherein X represents a carboxyl group, a sulfo group or a hydrogen atom, 1 and m each independently represent 0 or 1 and satisfy $1+m=1$, q represents 0 or 1, Z represents a hydrogen atom or a phenyl group that may be substituted with a carboxyl group or a sulfo group, provided that when q is 0, Z represents a phenyl group that may be substituted with a carboxyl group or a sulfo group.

[0006] The dyes disclosed in this publication show improvements in print density, light resistance and the like achieved by complicating functional groups of disazo dyes disclosed in Japanese Patent Laid-Open Publication No. H5(1993)-262998, etc. However, it is difficult for general synthetic methods by using γ acid (2-amino-8-naphthol-6-sulfonic acid) as a final coupling component to provide high purity.

[0007] The reason for this is that a derivative of γ acid intrinsically has a plurality of sites where the coupling reaction may take place (Chemistry of Synthetic Dyes, p. 154, published by Maki Shoten Co., Ltd.) and therefore, a single compound cannot be produced experimentally under alkaline condition through coupling with a diazonium salt made from a monoazo compound.

[0008] A dye generally contains a great large variety of impurities other than its main component and it is often difficult to analyze each of the impurity and to investigate its interaction with other components. However, as the amounts of the impurities increase, the probability that they interact mutually increases. The increase in the probability will appear as a degree of degradation, that is, as deterioration in storage stability or toughness such as light resistance when it is formulated into an ink.

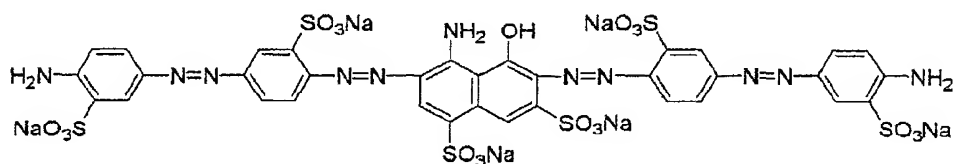
[0009] There are few dyes that have a structure similar to that of the trisazo compound of the present invention. For example, the dyes having structure of H acid (1-amino-8-naphthol-3,6-disulfonic acid) are known widely. Examples of the known compounds include Naphthol Blue Black of the structure shown below.



[0010] However, when an ink for ink jet recording is prepared by using only this disazo dye, optical density of the recorded image becomes insufficient, hue thereof also becomes unsatisfactory as a black ink, and its light resistance nor water resistance is not in the level for actual use.

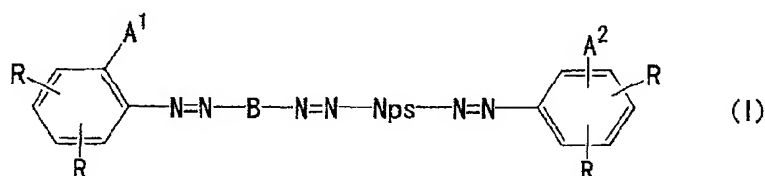
Further, the ink has a problem in dissolution stability of the dye.

[0011] In Japanese Patent Examined Publication No. S61(1986)-18590, a tetrakisazo dye of the structure shown below is used in a recording medium solution. However, it is, in general, difficult to obtain this kind of tetrakisazo dye as a high-purity product.

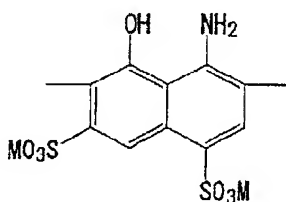


[0012] The objective of the present invention is to provide a high-purity trisazo compound with excellent light resistance that is useful as a coloring matter for a black aqueous ink for ink jet recording and the like.

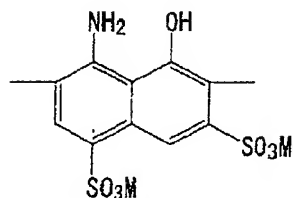
[0013] Thus, the present invention provides a trisazo compound represented by the formula:



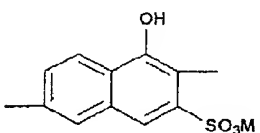
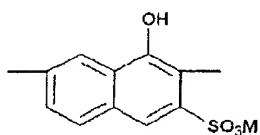
wherein A¹ is a group represented by the formula -COOM or -SO₃M wherein M is each independently a hydrogen atom, ammonium, alkali metal, hydroxyammonium, alkylammonium or alkanolammonium, A² is a hydrogen atom or a group represented by the formula -COOM or -SO₃M wherein M is the same as that described above, R is each independently a hydrogen atom, an alkyl group having from 1 to 6 carbon atoms, an alkoxy group having from 1 to 6 carbon atoms, a hydroxyl group, a sulfamoyl group, a carbamoyl group, a cyano group, an acetyl group, a sulfonic acid group, a carboxyl group or a halogen atom, B is a moiety represented by the formula:



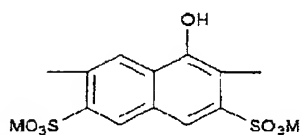
or



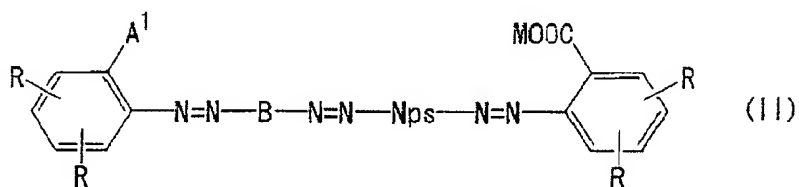
wherein, M is the same as that described above, Nps is a moiety represented by the formula:



or



wherein M is the same as that described above; or
the formula:



wherein M, A¹, R, B and Nps are the same as those described above.

[0014] Further, the present invention also provides a water-soluble dye containing the above-mentioned trisazo compound in an amount of not less than 85% by weight.

[0015] The water-soluble trisazo compound of the present invention has high purity. When the compound is used as a dye for an aqueous black ink for ink jet recording or the like, the ink has good dissolution stability and provides recorded images excellent in print density, light resistance and water resistance.

Fig. 1 shows a visible absorption spectrum of trisazo compound (1);

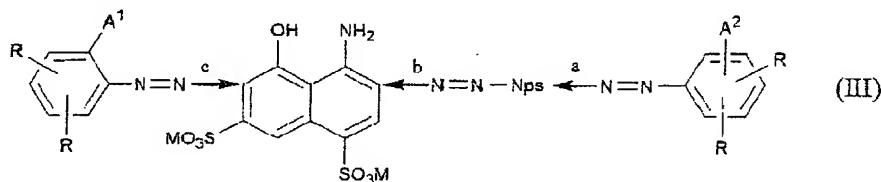
Fig. 2 shows a visible absorption spectrum of trisazo compound (2);

Fig. 3 shows a visible absorption spectrum of trisazo compound (5);

Fig. 4 shows a visible absorption spectrum of trisazo compound (6);
 Fig. 5 shows a visible absorption spectrum of trisazo compound (8);
 Fig. 6 shows a visible absorption spectrum of trisazo compound (10);
 Fig. 7 shows a visible absorption spectrum of trisazo compound (11);
 Fig. 8 shows an HPLC chart of trisazo compound (1);
 Fig. 9 shows an HPLC chart of trisazo compound (2);
 Fig. 10 shows an HPLC chart of trisazo compound (5);
 Fig. 11 shows an HPLC chart of trisazo compound (6);
 Fig. 12 shows an HPLC chart of trisazo compound (8);
 Fig. 13 shows an HPLC chart of trisazo compound (10);
 Fig. 14 shows an HPLC chart of trisazo compound (11);
 Fig. 15 shows an NMR spectrum of trisazo compound (1);
 Fig. 16 shows an NMR spectrum of trisazo compound (2);
 Fig. 17 shows an NMR spectrum of trisazo compound (5);
 Fig. 18 shows an NMR spectrum of trisazo compound (10);
 Fig. 19 shows an HPLC chart of trisazo compound (14);
 Fig. 20 shows a visible absorption spectrum of trisazo compound (14).

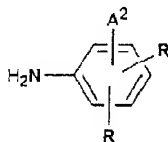
[0016] The trisazo compound (water-soluble dye) of the present invention can be produced via a diazotization-coupling process with reference to a well-known method, for example, a method described in Yutaka HOSODA, "New Dye Chemistry pp. 396 to 409", published by Gihodo Co., Ltd. on December 21, 1973.

[0017] For instance, the trisazo compound of the present invention can be obtained by subjecting a diazo compound to the coupling reaction in the order, a, b and c as shown in the following formula:

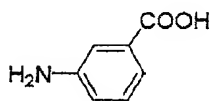
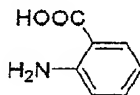


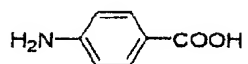
wherein M, A¹, A², R and Nps are the same as those described above. The synthesis procedure is described below.

[0018] In step a, an aniline derivative (1 mol) represented by formula:

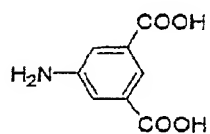


wherein A² and R are the same as those described above, is diazotized. Examples of the substituent R include a carboxyl group and an alkoxy group (for example, -OCH₃, and -OC₂H₅). Preferably A² is on ortho position with respect to an amino group. As the aniline derivatives, for example, compounds represented by the formulae:

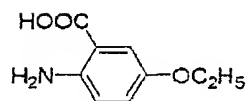




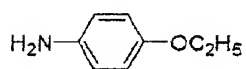
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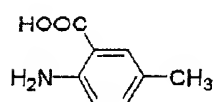
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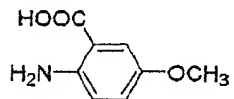
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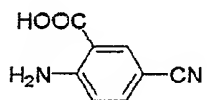
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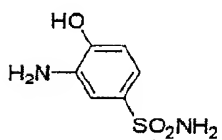
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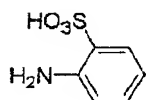
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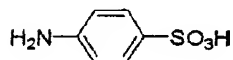
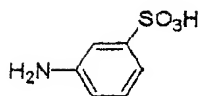


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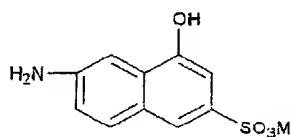
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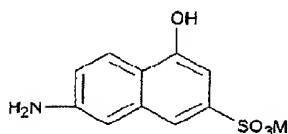


can be employed. Preferred aniline derivatives are anthranilic acid and phenetidine.

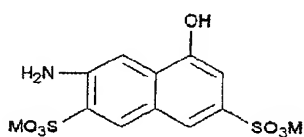
[0019] Next, the resulting diazotization product is subjected to alkaline coupling (pH=10.0 to 11.5) with a coupling component Nps-NH₂ (1 mol). Examples of Nps-NH₂ include γ acid represented by the formula:



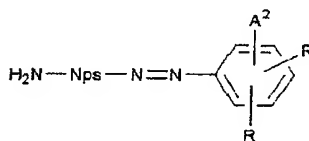
J acid represented by the formula:



and RR acid represented by the formula:

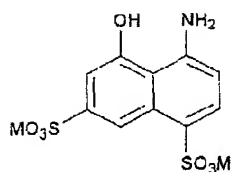


[0020] A preferred coupling component is γ acid. As a result, a monoazo intermediate shown below is obtained:

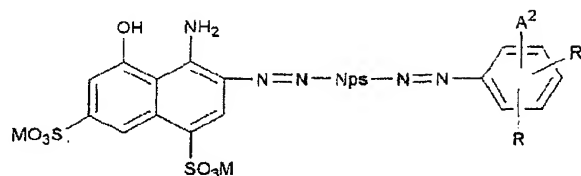


wherein A², R and Nps are the same as those described above.

[0021] In step b, the resulting monoazo intermediate (1 mol) is diazotized and then subjected to acidic coupling (pH not more than 4.0) with K acid (1 mol) represented by the formula:

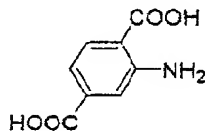
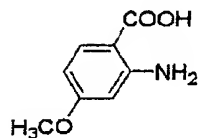
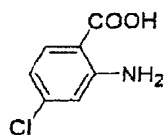
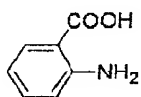


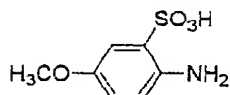
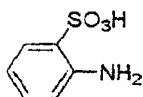
to obtain a disazo intermediate shown below:



wherein M, A², R and Nps are the same as those described above.

[0022] In step c, an aniline derivative (1 mol) having a carboxyl group or a sulfonic acid group on its ortho position is diazotized. It is desirable that the carboxyl group is converted to an ester group (e.g., -COOCH₃) prior to the diazotization reaction. If the reaction is carried out in the state where a free carboxyl group remains, yield or purity of the compound may become poor. Examples of the aniline derivative include compounds represented by the formulae:

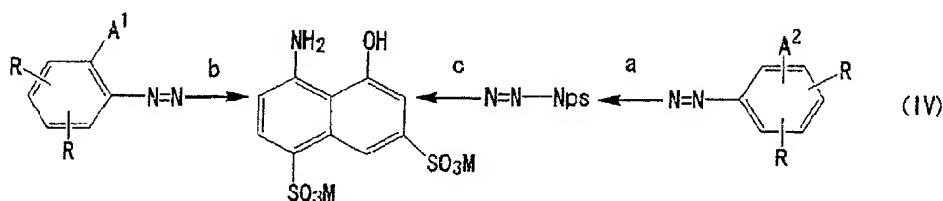




[0023] A preferred aniline derivative is anthranilic acid.

[0024] Subsequently, the resulting diazotization product is subjected to alkaline coupling (pH=9.5 to 11.5) with the above-mentioned disazo intermediate (1 mol) as a coupling component and, if an ester group is contained, the ester group is converted to a carboxyl group by hydrolysis, followed by desaltation and purification, to obtain the trisazo compound of the present invention represented by formula (III).

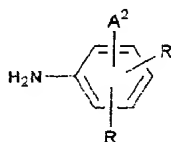
[0025] Otherwise, the trisazo compound of the present invention can be obtained by subjecting a diazo compound to the coupling reaction in the order, a, b and c as shown in the following formula:



wherein A^1 , A^2 , R, Nps and M are the same as those described above. The compounds and conditions for the coupling reaction are the same as those described above in the synthesis procedure for the trisazo compound (III).

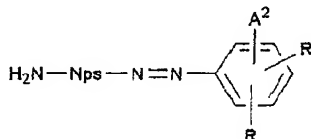
[0026] More specifically in this instance, the trisazo compound of the present invention represented by formula (IV) may be produced by the synthesis procedure which comprises the steps of:

a1) diazotizing an aniline derivative represented by the formula:



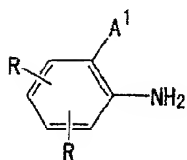
wherein, A^2 and R are the same as those described above;

a2) subjecting the resulted diazotization product to alkaline coupling with a coupling component selected from the group consisting of γ acid, J acid and RR acid to obtain a monoazo intermediate represented by the formula:



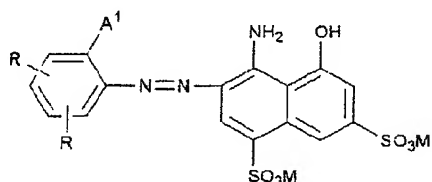
wherein A^2 and R are the same as those described above, and Nps is a residue of γ acid, J acid, or RR acid;

b1) diazotizing an aniline derivative having a carboxyl group or a sulfonic group on its ortho position represented by the formula:



wherein A¹ and R are the same as those described above; and

b2) subjecting the resulted diazotization product to acidic coupling with K acid to obtain a monoazo intermediate represented by the formula:

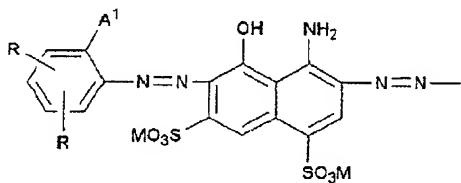


wherein A¹, R and M are the same as those described above;

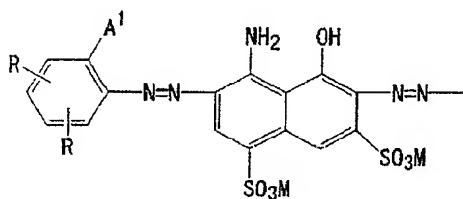
c1) diazotizing the monoazo intermediate resulted in step a2); and

c2) subjecting the resulted diazotization product to alkaline coupling with the monoazo intermediate resulted in step b2).

[0027] As described above, K acid is employed for a raw material in synthesis of the trisazo compound of the present invention. By using K acid, structure specific to the trisazo compound of the present invention represented by the formula:



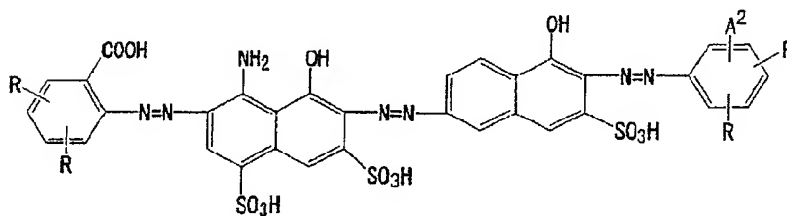
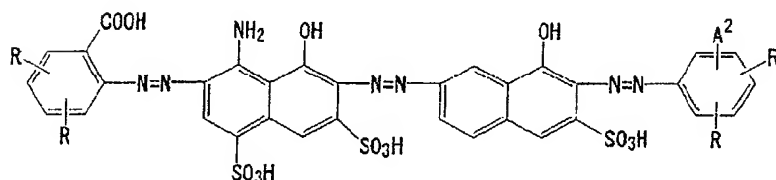
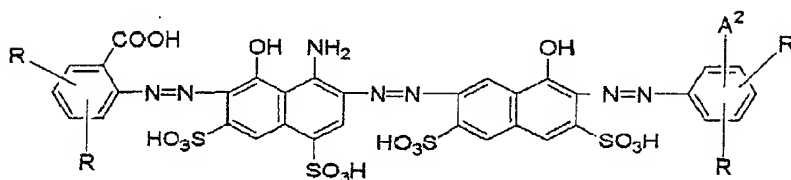
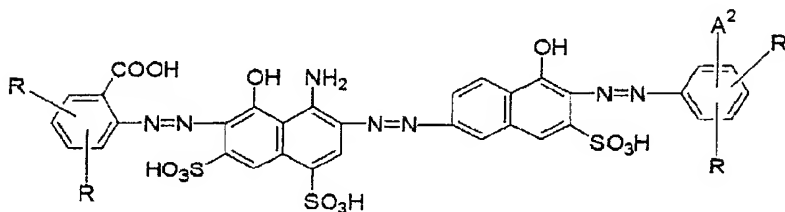
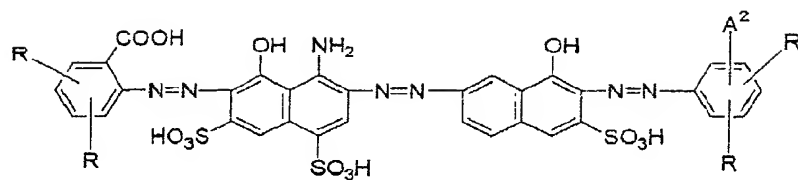
or



wherein M, A¹ and R are the same as those described above, is formed.

[0028] The trisazo compound (dye) of the present invention obtained in such manner is a high-purity product containing very small amount of impurities and has a purity determined by HPLC (high performance liquid chromatography) of preferably not less than about 85%, more preferably not less than 90%.

[0029] Preferred examples of the trisazo compound of the present invention are shown in the form of free acids as follows:



wherein A^2 is a hydrogen atom, a carboxyl group or a sulfonic group, R is each independently a hydrogen atom, an alkyl group having 1 to 6 carbon atoms, an alkoxy group having 1 to 6 carbon atoms, a hydroxyl group, a sulfamoyl group, a carbamoyl group, a cyano group, an acetyl group, a sulfonic group, a carboxyl group or a halogen atom.

[0030] Specific examples of the trisazo compound of the present invention in the form of free acid are shown below together with their maximum absorption wavelength (λ_{\max}). However, this invention is not limited thereto.

Table 1

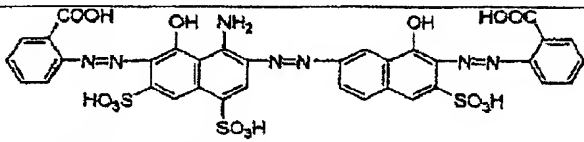
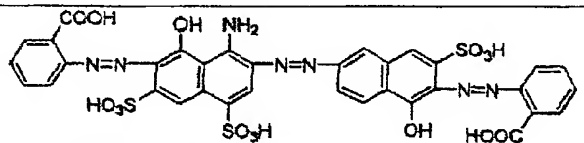
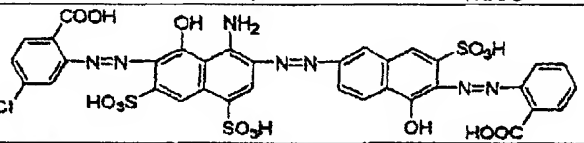
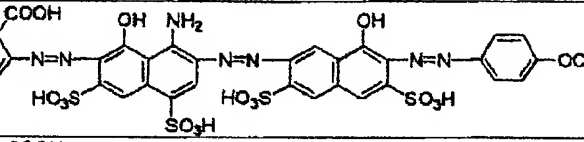
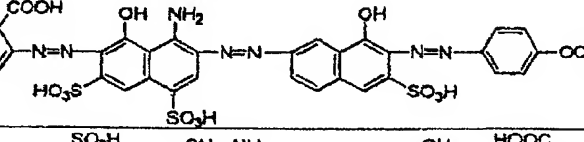
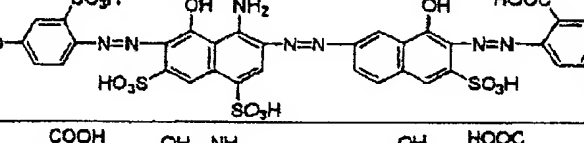
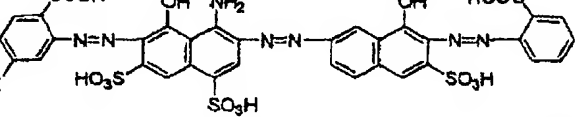
	Examples of trisazo compound	λ_{\max} (nm)
(1) AN-250		603.2 512.4
(2) AN-251		608.0
(3) AN-252		601.1
(4) AN-253		605.2
(5) AN-254		608.0
(6) AN-255		614.2 507.8
(7) AN-256		604.6

Table 2

	Examples of trisazo compound	λ_{\max} (nm)
(8) AN-257		597.6
(9) AN-258		608.4
(10) AN-259		604.2
(11) AN-260		597.6
(12)		601.4
(13)		596.2

Table 3

	Examples of trisazo compound	λ_{\max} (nm)
(14) AN-261		624.4
(15)		626.0
(16)		621.2

[0031] The trisazo compound of the present invention exhibits high print density and high fastness (fastness with respect to light resistance, water resistance and stability) when it is used as a dye for an ink for ink jet recording. The cause of this has not been fully investigated, but it is probably because the compound has a structure with a combination of one aromatic ring having a carboxyl group or a sulfonic group on its ortho position with respect to a diazo bond and another aromatic ring derived from K acid. That is, the arrangement of the functional groups makes it difficult to form impurities during the coupling synthesis sequentially carried out and, as a result, oxide or the like that will deteriorate toughness is not formed, and the product becomes very high in purity.

[0032] In the present specification, although the specific examples of water-soluble groups (sulfonic groups and carboxyl groups) of trisazo compounds are expressed in the form of free acids, the compounds are dissolved in water in the form of salts together with alkali metals or amines. Examples of the alkali metals include Na, K and Li. Examples of the amines include ammonia; hydroxylamine; alkyl-substituted amines such as methylamine, dimethylamine, trimethylamine, ethylamine, diethylamine, triethylamine, propylamine, isopropylamine, dipropylamine, diisopropylamine, butylamine, isobutylamine, sec-butylamine, tert-butylamine, dibutylamine and allylamine; alkanolamines such as ethanolamine, diethanolamine, triethanolamine, propanolamine and diethylethanolamine; and their mixtures.

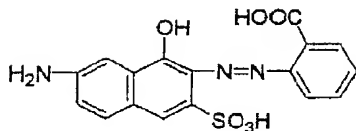
[0033] The following Examples further illustrates the present invention in detail but are not to be construed to limit the scope thereof.

Example 1

Synthesis example of trisazo compound (1) (AN-250)

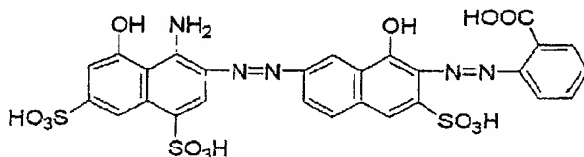
[0034] Synthesis of monoazo intermediate: To 150 ml of water, 27.4 g (0.20 mol) of anthranilic acid was charged. This was dissolved by adding 20%-NaOH aqueous solution until the pH became 7.0. To the resulting solution, 42.2 g (0.22 mol) of a 36%-NaNO₃ aqueous solution was added and the temperature was lowered to 0°C by adding ice. The pH was thereafter adjusted to 0.8 by adding 69.0 g (0.67 mol) of 35%-HCl aqueous solution. After stirring at 4°C for 50 minutes, a diazo solution was obtained.

[0035] To 450 ml of water, on the other hand, 47.9 g (0.20 mol) of γ acid was charged. This was dissolved by adding 20%-NaOH aqueous solution until the pH became 7.0. The temperature of the solution was lowered to 0°C in a cooling bath. A coupling reaction was carried out by gradually adding the diazo solution prepared previously to the cooled solution. During the reaction, the pH of the reaction solution was regulated between 10.0 and 11.5 by adding 20%-NaOH aqueous solution and, at the same time, the reaction temperature was held so as not to exceed 7°C. After stirring the reaction solution at 6°C overnight, the pH of the solution was adjusted to 3.3 by dropping 35%-HCl aqueous solution. The precipitate was collected by filtration to obtain the monoazo intermediate shown below (purity: 95.2% determined by HPLC).



[0036] Synthesis of disazo intermediate: To 1000 ml of water, a press cake of the monoazo intermediate (about 0.2 mol) was charged. This was dissolved by adding 20%-NaOH aqueous solution until the pH became 6.1. To the resulting solution, 42.2 g (0.22 mol) of 36%-NaOH aqueous solution was added and it was cooled to 15°C. The pH was thereafter adjusted to 0.8 by adding 102.4 g (1.0 mol) of 35%-HCl aqueous solution. After stirring at 23°C for 70 minutes, a diazo solution was obtained.

[0037] To 600 ml of water, on the other hand, 68.3 g (0.20 mol) of K acid was charged. This was dissolved by adding sodium carbonate. A coupling reaction was carried out under acidic condition by dropping the diazo solution prepared previously to the solution. During the reaction, the reaction temperature was regulated so as not to exceed 6°C. After the addition of the diazo solution, stirring was continued for 1 hour and then additional stirring was carried out overnight while adjusting the pH to not more than 4.0 by dropping 10%-Na₂CO₃ aqueous solution. The precipitate was collected by filtration to obtain a disazo intermediate shown below (purity: 80.3% determined by HPLC).



[0038] Synthesis of trisazo compound: To 150 ml of water, 28.7 g (0.19 mol) of methyl anthranilate was added and then 150 g of ice was charged. Thereafter, 51.4 g (0.50 mol) of 35%-HCl aqueous solution was added. After cooling this mixture to 0°C or lower, 40.3 g (0.21 mol) of 36%-NaNO₃ aqueous solution was added and it was stirred at 3°C for 50 minutes to obtain a diazo solution.

[0039] Separately, a press cake of the disazo intermediate (about 0.19 mol) was charged to 1700 ml of water this was dissolved by adding 20%-NaOH aqueous solution until the pH became 11.0. The temperature was lowered to -1°C in a cooling bath. To the mixture, the diazo solution prepared previously was added gradually to carry out coupling. During this operation, the reaction condition was adjusted so that the pH of the reaction solution became from 8.0 to 11.0 by adding 20%-NaOH aqueous solution and the reaction temperature was held so as not to exceed 7°C. After stirring the reaction solution at 6°C overnight, 680 g of table salt was added thereto and was stirred for a short period of time. A precipitate was collected by filtration.

[0040] The press cake obtained was charged to 1500 ml of water and the pH was adjusted to 12.5 or higher by adding 20%-NaOH aqueous solution and the mixture was stirred at 60°C for 80 minutes. When the mixture was cooled to 40°C, a precipitate was collected by filtration. The press cake obtained was dissolved again in a proper amount of water. The solution was desalted with a reverse osmosis membrane (NTR-7430 manufactured by Nippon Denko Co., Ltd.) and then the dye desalted solution was dried by freeze drying to obtain objective trisazo compound (1), AN-250 (purity : 96.3% determined by HPLC, λ_{max} : 603.2 nm (in water)).

Table 4

Analysis data	
Visible absorption spectrum	Fig. 1
HPLC (High-Performance Liquid Chromatography)	Fig. 8
NMR spectrum	Fig. 15

Evaluation of light resistance

Preparation of samples for test:

[0041]

Table 5

Trisazo compound (dye) of the present invention	6.0% by weight
Glycerin	10% by weight
Diethyleneglycol monobutyl ether	10% by weight
Water	The remainder

[0042] An ink, the pH of which was adjusted between 9.0 and 9.5 with ammonia, prepared using the composition shown above was set into an ink jet printer (EM-900C manufactured by Seiko Epson Corp.). A piece of paper specially designed for ink jet (PM Photo Paper (Model: KA420PSK) manufactured by Seiko Epson Corp.) was printed with the ink, and it was used as a sample to be tested.

[0043] Light irradiation: The sample to be tested was set to a holder. Light irradiation with a xenon fade meter was carried out under the following condition:

Table 6

Instrument used	ATLAS Ci4000 Xenon Weather-Ometer
Light source filter	Borosilicate/Soda lime
Intensity of irradiation	36 W/m ² (300 to 400 nm)
Chamber temperature (Temperature in chamber)	24°C
Chamber humidity	55% j

[0044] Method for evaluation: The sample is taken out after each irradiation of an energy of 3000 kJ/m², and a part with an optical density (OD value) of about 1.0 is measured with a Macbeth density meter. The total amount of light irradiated at the time when the OD value has become 70% the original value is calculated.

[0045] Criterion for evaluation: The amount of energy required for the OD value to decrease to 70%.

Table 7

⊙	Not less than 70,000 kJ/m ²
○	Not less than 40,000 kJ/m ² and less than 70,000 kJ/m ²
Δ	Not less than 10,000 kJ/m ² and less than 40,000 kJ/m ²
×	Not more than 10,000 kJ/m ²

[0046] Trisazo compound (dye) (1) has a solubility in water (under an ammoniacal alkaline condition) of 14%. The result of its mutagenicity test (Ames test) proved negative (1.27 times). The results of the evaluation are shown in Table 15.

Example 2

Synthesis example of trisazo compound (2) (AN-251)

[0047] Trisazo compound (2) was obtained in substantially the same manner as described in Example 1, except that J acid (3-amino-8-naphthol-6-sulfonic acid) was used instead of γ acid used in Example 1, and the resulting compound was evaluated. The results of the evaluation are shown in Table 15.

Table 8

Analysis data	
Visible absorption spectrum	Fig. 2
HPLC (High-Performance Liquid Chromatography)	Fig. 9
NMR spectrum	Fig. 16

Example 3

Synthesis example of trisazo compound (5) (AN-254)

[0048] Trisazo compound (5) was obtained in substantially the same manner as described in Example 1, except that p-phenetidine was used instead of the anthranilic acid used in Example 1, and was evaluated. The results of the evaluation are shown in Table 15.

Table 9

Analysis data	
Visible absorption spectrum	Fig. 3
HPLC (High-Performance Liquid Chromatography)	Fig. 10
NMR spectrum	Fig. 17

Example 4

Synthesis example of trisazo compound (6) (AN-255)

[0049] Trisazo compound (6) was obtained in substantially the same manner as described in Example 1, except that 4-chloro-anthranilic acid and p-anisidine-2-sulfonic acid were used instead of the anthranilic acid and the methyl anthranilate used in Example 1 respectively, and the resulting compound was evaluated. The results of the evaluation are shown in Table 15.

Table 10

Analysis data	
Visible absorption spectrum	Fig. 4
HPLC (High-Performance Liquid Chromatography)	Fig. 11

Example 5

Synthesis example of trisazo compound (8) (AN-257)

[0050] Trisazo compound (8) was obtained in substantially the same manner as described in Example 1, except that 2-aminoterephthalic acid was used instead of the methyl anthranilate used in Example 1, and the resulting compound was evaluated. The results of the evaluation are shown in Table 15.

Table 11

Analysis data	
Visible absorption spectrum	Fig. 5
HPLC (High-Performance Liquid Chromatography)	Fig. 12

Example 6

Synthesis example of trisazo compound (10) (AN-259)

- 5 **[0051]** Trisazo compound (10) was obtained in substantially the same manner as described in Example 1, except that p-n-butylaniline was used instead of the anthranilic acid used in Example 1, and the resulting compound was evaluated. The results of the evaluation are shown in Table 15.

Table 12

10	Analysis data	
	Visible absorption spectrum	Fig. 6
	HPLC (High-Performance Liquid Chromatography)	Fig. 13
15	NMR spectrum	Fig. 18

Example 7

Synthesis example of trisazo compound (11) (AN-260)

- 20 **[0052]** Trisazo compound (11) was obtained in substantially the same manner as described in Example 1, except that 3-aminobenzoic acid was used instead of the anthranilic acid used in Example 1, and the resulting compound was evaluated. The results of the evaluation are shown in Table 15.

Table 13

25	Analysis data	
	Visible absorption spectrum	Fig. 7
	HPLC (High-Performance Liquid Chromatography)	Fig. 14
30		

Example 8

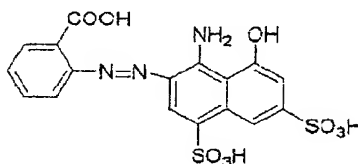
Synthesis example of trisazo compound (14) (AN-261)

- 35 1) Synthesis of monoazo intermediate (1)

[0053] To 300 ml of water, 67.95 g (0.3 mol) of methyl anthranilate was charged. To this, 150 g of ice and 123.0 g (1.2 mol) of 35%-HCl aqueous solution was further added. This was cooled to 0°C, 86.25 g (0.45 mol) of 36%-NaNO₃ aqueous solution was added, and stirred at 3°C for 30 minutes to obtain a diazo solution.

40 **[0054]** On the other hand, 121.2 g (0.3 mol) of K acid was charged to 1000 ml of water. This was dissolved by adding sodium carbonate. A coupling reaction was carried out under acidic condition by dropping the diazo solution prepared previously to the resulting solution. During the reaction, the reaction temperature was regulated so as not to exceed 6°C. After the addition of the diazo solution, stirring was continued for 1 hour and then additional stirring was carried out while adjusting the pH to 1.5 by dropping 10%-Na₂CO₃ aqueous solution. A precipitate was collected by filtration.

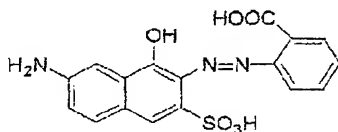
45 **[0055]** The resulting pressed cake was charged to 1000 ml of water, pH thereof was adjusted to not less than 12.5 by adding 20%-NaOH aqueous solution, and the mixture was stirred for 90 minutes at 60°C. This was cooled to 40°C, 260 g of salt was added, stirred for a while, pH thereof was adjusted to 3.3 by adding 35%-HCl aqueous solution, and a precipitate was collected by filtration to obtain the monoazo intermediate as shown below (purity: 88.4% determined by HPLC).



2) Synthesis of monoazo intermediate (2)

[0056] To 300 ml of water, 41.1 g (0.3 mol) of anthranilic acid was charged. This was dissolved by adding 20%-NaOH aqueous solution until the pH became 7.0. To the resulting solution, 63.3 g (0.33 mol) of a 36%-NaNO₃ aqueous solution was added and the temperature was lowered to 0°C by adding ice. The pH was thereafter adjusted to 0.8 by adding 112.1 g (1.09 mol) of 35%-HCl aqueous solution. After stirring at 3°C for 55 minutes, a diazo solution was obtained.

[0057] To 600 ml of water, on the other hand, 73.8 g (0.3 mol) of γ acid was charged and 20%-NaOH aqueous solution was added until the pH became 9. The temperature of the solution was lowered to 0°C in a cooling bath. A coupling reaction was carried out by gradually adding the diazo solution prepared previously to the cooled solution. During the reaction, the pH of the reaction solution was regulated between 10.0 and 11.5 by adding 20%-NaOH aqueous solution and, at the same time, the reaction temperature was held so as not to exceed 7°C. After stirring the reaction solution at 10°C overnight, the pH of the solution was adjusted to 4.0 by dropping 35%-HCl aqueous solution. A precipitate was collected by filtration to obtain the monoazo intermediate as shown below (purity: 98.5% determined by HPLC).



3) Synthesis of trisazo compound

[0058] To 1000 ml of water, a pressed cake of the monoazo intermediate (0.1 mol) obtained by step 2) was charged. This was dissolved by adding 20%-NaOH aqueous solution until pH became 7.4. To the resulting solution, 23 g (0.12 mol) of 36%-NaNO₃ aqueous solution was added and it was cooled to 18°C. The pH was thereafter adjusted to 0.7 by adding 61.4 g (0.6 mol) of 35%-HCl aqueous solution. After stirring at 20°C for 90 minutes, a diazo solution was obtained.

[0059] On the other hand, a pressed cake of the monoazo intermediate (0.1 mol) was charged to 700 ml of water. This was dissolved by adding 20%-NaOH aqueous solution until pH became 10 and the temperature was lowered to 0°C in an ice bath. A coupling reaction was carried out by gradually adding the diazo solution prepared previously to the cooled solution. During the reaction, the pH was regulated so as to be 8.0 to 11.0 by adding 20%-NaOH aqueous solution, and the reaction temperature was held so as not to exceed 7°C. After stirring the reaction solution at 12°C overnight, 250 g of table salt was added thereto and was stirred for a short period of time. A precipitate was collected by filtration.

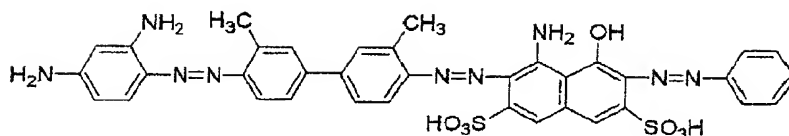
[0060] After the trisazo dye obtained above was dissolved in 1200 ml of water, 35%-HCl aqueous solution was added so that the pH became 1 to 2, and the dye was isolated by filtration. The press cake obtained was dissolved again in a proper amount of water by using aqueous ammonia. The solution was desalted with a reverse osmosis membrane (NTR-7430 manufactured by Nitto Denko Co., Ltd.). The dye desalted solution was then dried by freeze drying to obtain objective trisazo compound (14), AN-261 (purity : 94.4% determined by HPLC, λ_{\max} : 624.4 nm (in water)). The resulted trisazo compound (14) was evaluated according to the same manner as described in Example 1. The results are shown in Table 15.

Table 14

Analysis data	
Visible absorption spectrum	Fig. 20
HPLC (High-Performance Liquid Chromatography)	Fig. 19

Comparative Example 1

[0061] C. I. Direct Black 154 was used as dye 1 for comparison. The results of the evaluation are shown in Table 15.



Comparative Example 2

[0062] C. I. Direct Black 168 was used as dye 2 for comparison. The results of the evaluation are shown in Table 15.

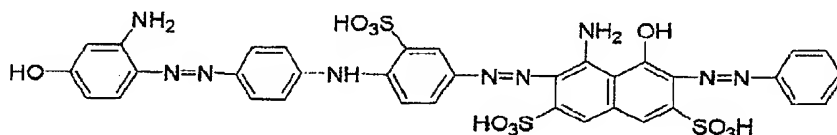
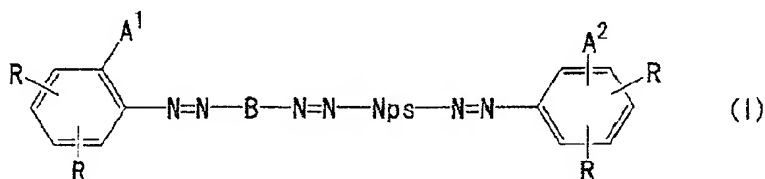


Table 15

	Example of dye (salt)	Purity by HPLC (%)	Evaluation of light resistance
Example 1	Trisazo dye (1) (NH ₄)	96.3	⊙
Example 2	Trisazo dye (2) (NH ₄)	97.6	⊙
Example 3	Trisazo dye (5) (NH ₄)	96.4	⊙
Example 4	Trisazo dye (6) (NH ₄)	95.6	○
Example 5	Trisazo dye (8) (NH ₄)	86.1	⊙
Example 6	Trisazo dye (10) (NH ₄)	92.9	○
Example 7	Trisazo dye (11) (NH ₄)	84.4	○
Example 8	Trisazo dye (14) (NH ₄)	94.4	⊙
Comparative Example 1	C. I. Direct Black 154 (NH ₄)	-	×
Comparative Example 2	C. I. Direct Black 168 (NH ₄)	-	Δ

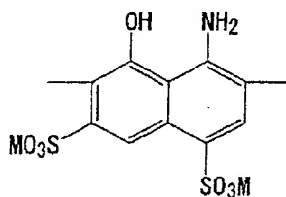
Claims

1. A trisazo compound represented by the formula:

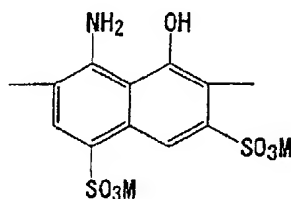


wherein A¹ is a group represented by the formula -COOM or -SO₃M wherein M is each independently a hydrogen atom, ammonium, alkali metal, hydroxyammonium, alkylammonium or alkanolammonium, A² is a hydrogen atom or a group represented by the formula -COOM or -SO₃M wherein M is the same as that described above, R is

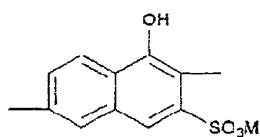
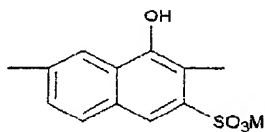
each independently a hydrogen atom, an alkyl group having from 1 to 6 carbon atoms, an alkoxy group having from 1 to 6 carbon atoms, a hydroxyl group, a sulfamoyl group, a carbamoyl group, a cyano group, an acetyl group, a sulfonic group, a carboxyl group or a halogen atom, B is a moiety represented by the formula:



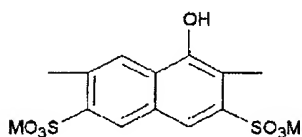
or



wherein, M is the same as that described above, Nps is a moiety represented by the formula:

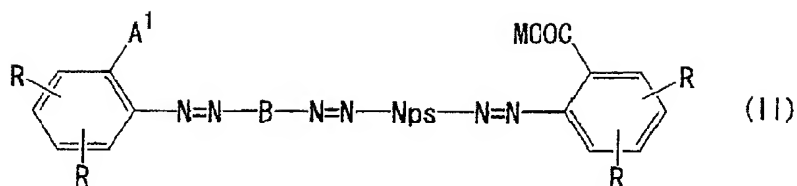


or



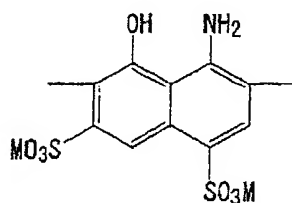
wherein M is the same as that described above.

2. A trisazo compound represented by the formula;

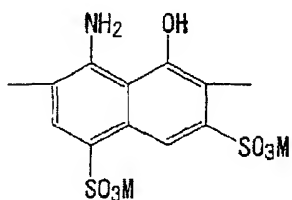


10 wherein M is each independently a hydrogen atom, ammonium, alkali metal, hydroxyammonium, alkylammonium or alkanolammonium, A¹ is a group represented by the formula -COOM or -SO₃M wherein M is the same as that described above, R is each independently a hydrogen atom, an alkyl group having from 1 to 6 carbon atoms, an alkoxy group having from 1 to 6 carbon atoms, a hydroxyl group, a sulfamoyl group, a carbamoyl group, a cyano group, an acetyl group, a sulfonic group, a carboxyl group or a halogen atom, B is a moiety represented by the formula:

15

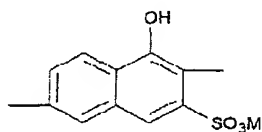
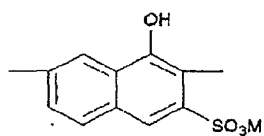


25 or

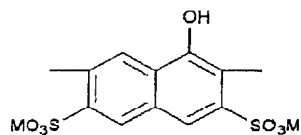


wherein, M is the same as that described above, Nps is a moiety represented by the formula:

40

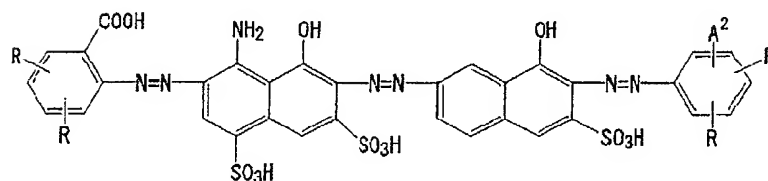
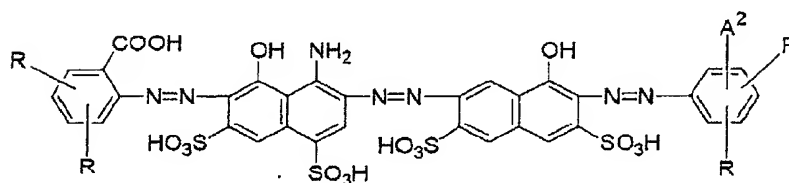
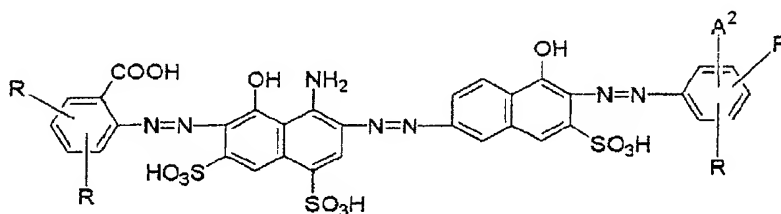
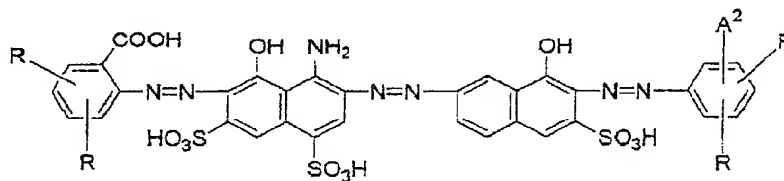


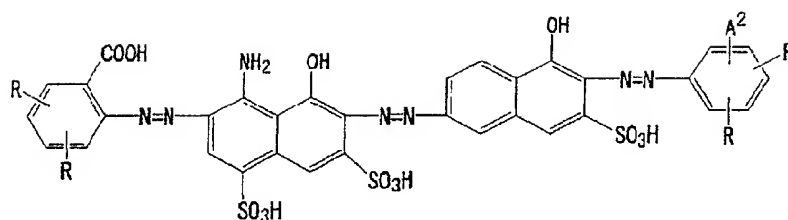
or



wherein M is the same as that described above.

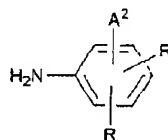
3. The trisazo compound according to claim 1 or 2, wherein said A¹ is a group represented by the formula -COOM wherein M is the same as that described in claim 1 or 2.
4. A trisazo compound having a structure represented by any one of the following formulae in the form of free acids:





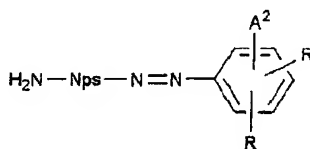
wherein A^2 is a hydrogen atom, a carboxyl group or a sulfonic group, R is each independently a hydrogen atom, an alkyl group having 1 to 6 carbon atoms, an alkoxy group having 1 to 6 carbon atoms, a hydroxyl group, a sulfamoyl group, a carbamoyl group, a cyano group, an acetyl group, a sulfonic group, a carboxyl group or a halogen atom.

5. A water-soluble dye which comprises the trisazo compound of any one of claims 1 to 4 in an amount of not less than 85% by weight.
6. A process for producing the trisazo compound of claim 1 comprising the steps of:
 - a1) diazotizing an aniline derivative represented by the formula:



wherein, A^2 is a hydrogen atom or a group represented by the formula -COOM or $-SO_3M$ wherein M is each independently a hydrogen atom, ammonium, alkali metal, hydroxyammonium, alkylammonium or alkanolammonium, R is each independently a hydrogen atom, an alkyl group having from 1 to 6 carbon atoms, an alkoxy group having from 1 to 6 carbon atoms, a hydroxyl group, a sulfamoyl group, a carbamoyl group, a cyano group, an acetyl group, a sulfonic group, a carboxyl group or a halogen atom;

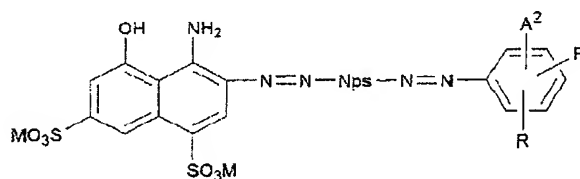
a2) subjecting the resulted diazotization product to alkaline coupling with a coupling component selected from the group consisting of γ acid, J acid and RR acid to obtain a monoazo intermediate represented by the formula:



wherein A^2 and R are the same as those described above, and Nps is a residue of γ acid, J acid, or RR acid;

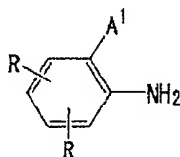
b1) diazotizing the monoazo intermediate;

b2) subjecting the resulted diazotization product to acidic coupling with K acid to obtain a disazo intermediate represented by the formula:



wherein M, A², R and Nps are the same as those described above;

c1) diazotizing an aniline derivative having a carboxyl group or a sulfonic group on its ortho position represented by the formula:

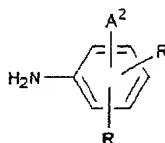


wherein A¹ is a group represented by the formula -COOM or -SO₃M wherein M is the same as that described above, and R is the same as that described above; and

c2) subjecting the resulted diazotization product to alkaline coupling with the disazo intermediate resulted in step b2).

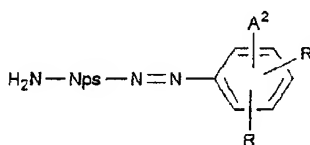
7. A process for producing the trisazo compound of claim 1 comprising the steps of:

a1) diazotizing an aniline derivative represented by the formula:

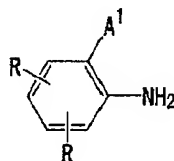


wherein, A² is a hydrogen atom or a group represented by the formula -COOM or -SO₃M wherein M is each independently a hydrogen atom, ammonium, alkali metal, hydroxy ammonium, alkylammonium or alkanolammonium, R is each independently a hydrogen atom, an alkyl group having from 1 to 6 carbon atoms, an alkoxy group having from 1 to 6 carbon atoms, a hydroxyl group, a sulfamoyl group, a carbamoyl group, a cyano group, an acetyl group, a sulfonic group, a carboxyl group or a halogen atom;

a2) subjecting the resulted diazotization product to alkaline coupling with a coupling component selected from the group consisting of γ acid, J acid and RR acid to obtain a monoazo intermediate represented by the formula:

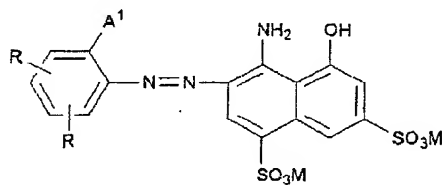


wherein A² and R are the same as those described above, and Nps is a residue of γ acid, J acid, or RR acid; b1) diazotizing an aniline derivative having a carboxyl group or a sulfonic group on its ortho position represented by the formula:



wherein A¹ is a group represented by the formula -COOM or -SO₃M wherein M is the same as that described

above, and R is the same as that described above; and
 b2) subjecting the resulted diazotization product to acidic coupling with K acid to obtain a monoazo intermediate represented by the formula:



wherein A¹, R and M are the same as those described above;
 c1 diazotizing the monoazo intermediate resulted in step a2); and
 c2) subjecting the resulted diazotization product to alkaline coupling with the monoazo intermediate resulted in step b2).

Fig. 1

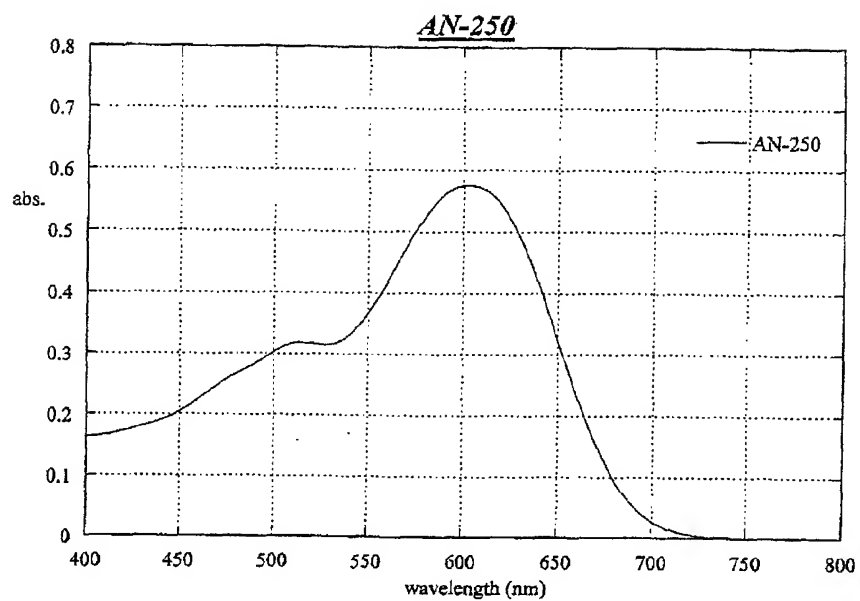


Fig. 2

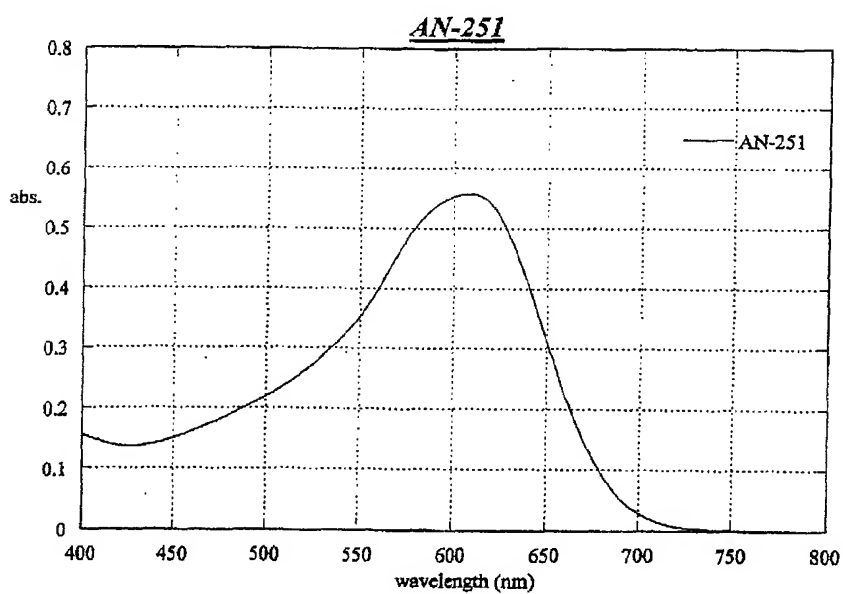


Fig. 3

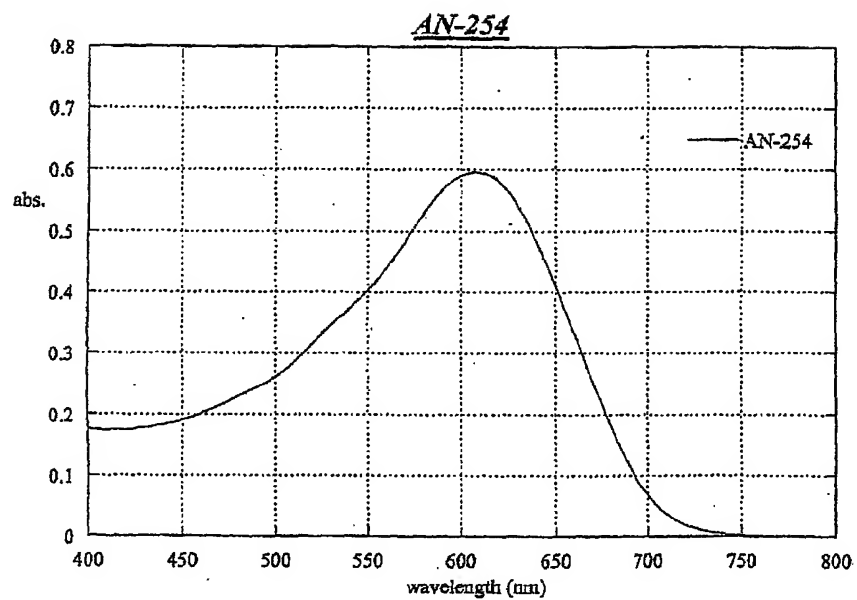


Fig. 4

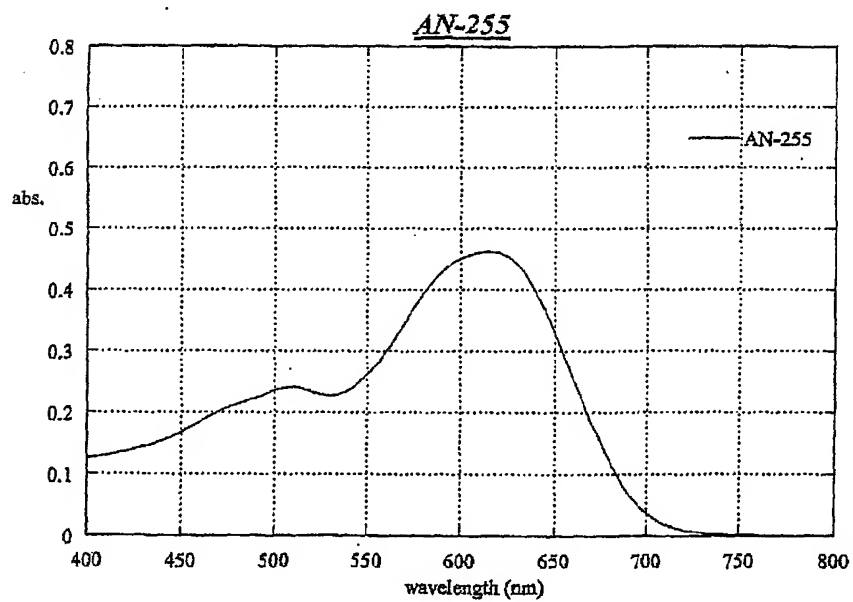


Fig. 5

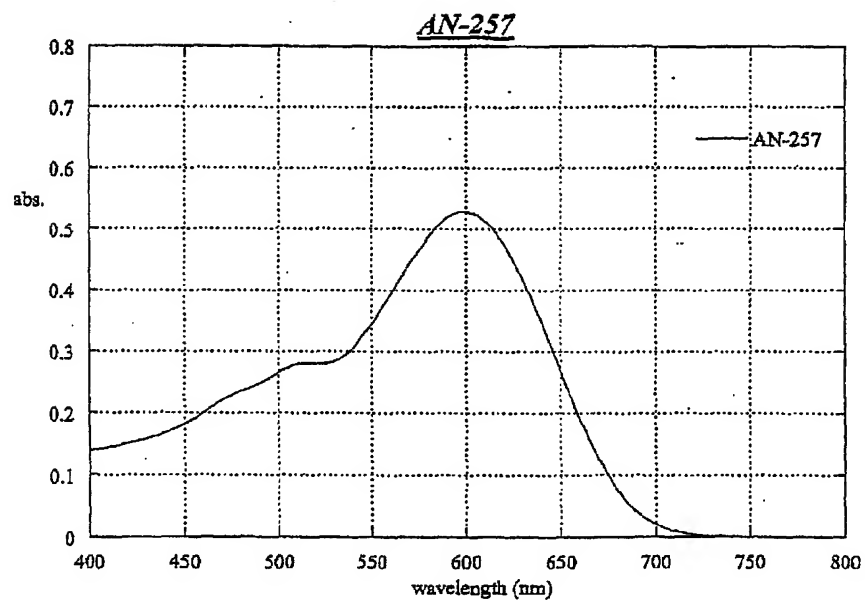


Fig. 6

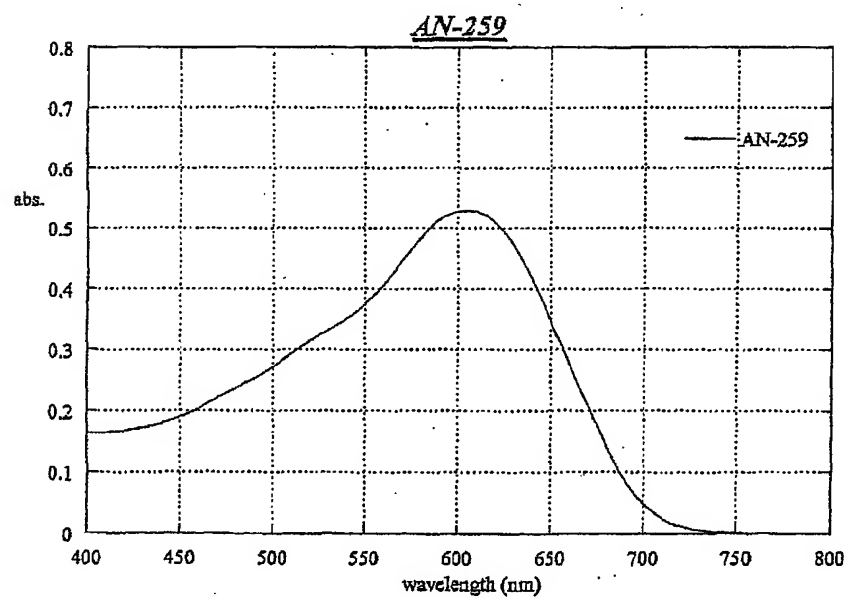


Fig. 7

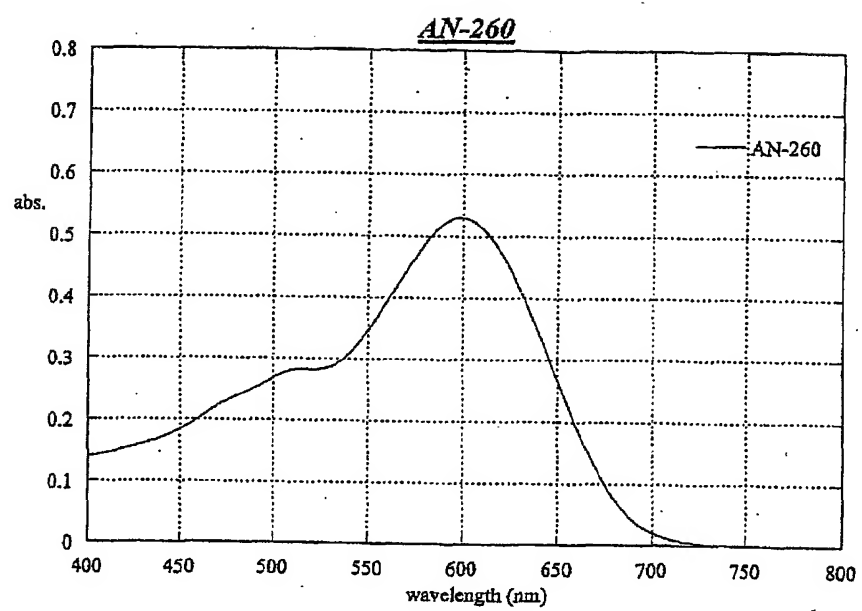


Fig. 8

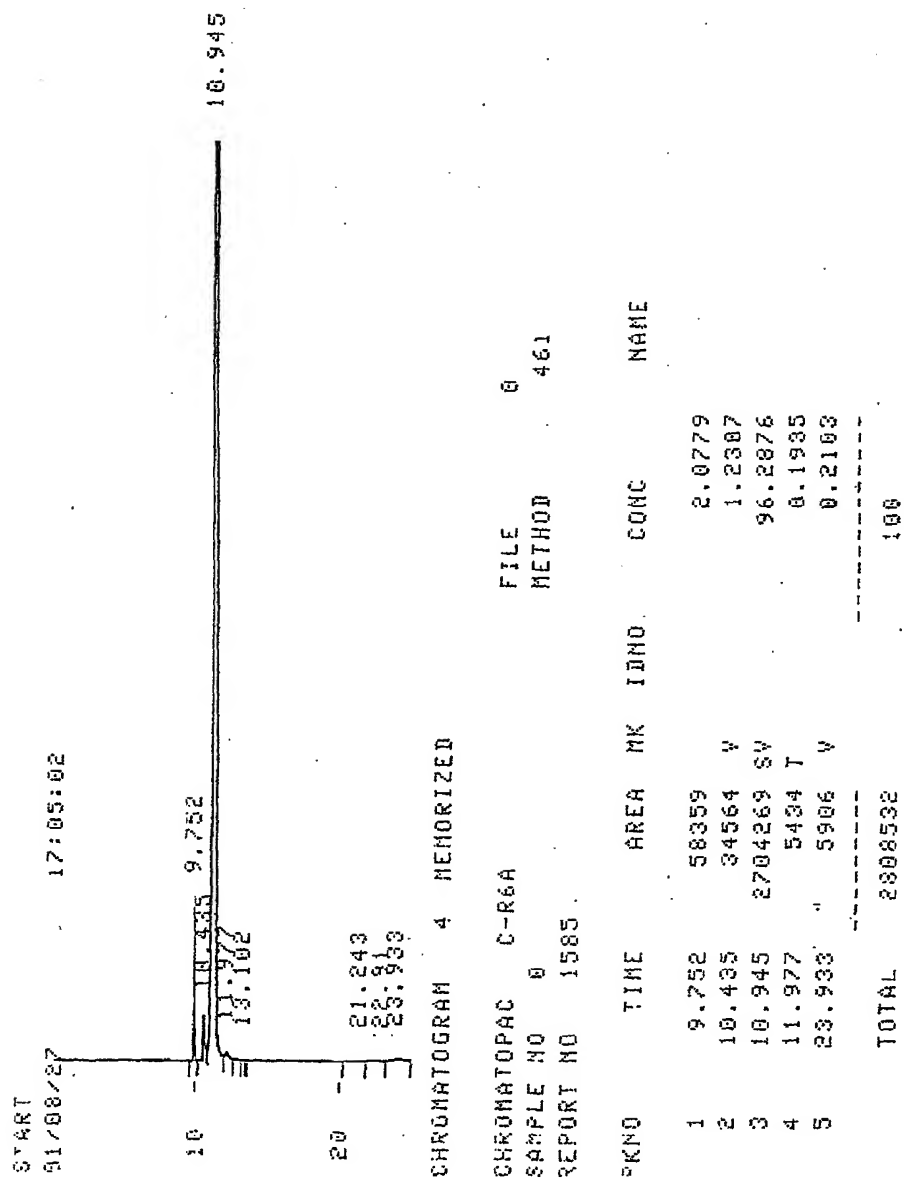


Fig. 9

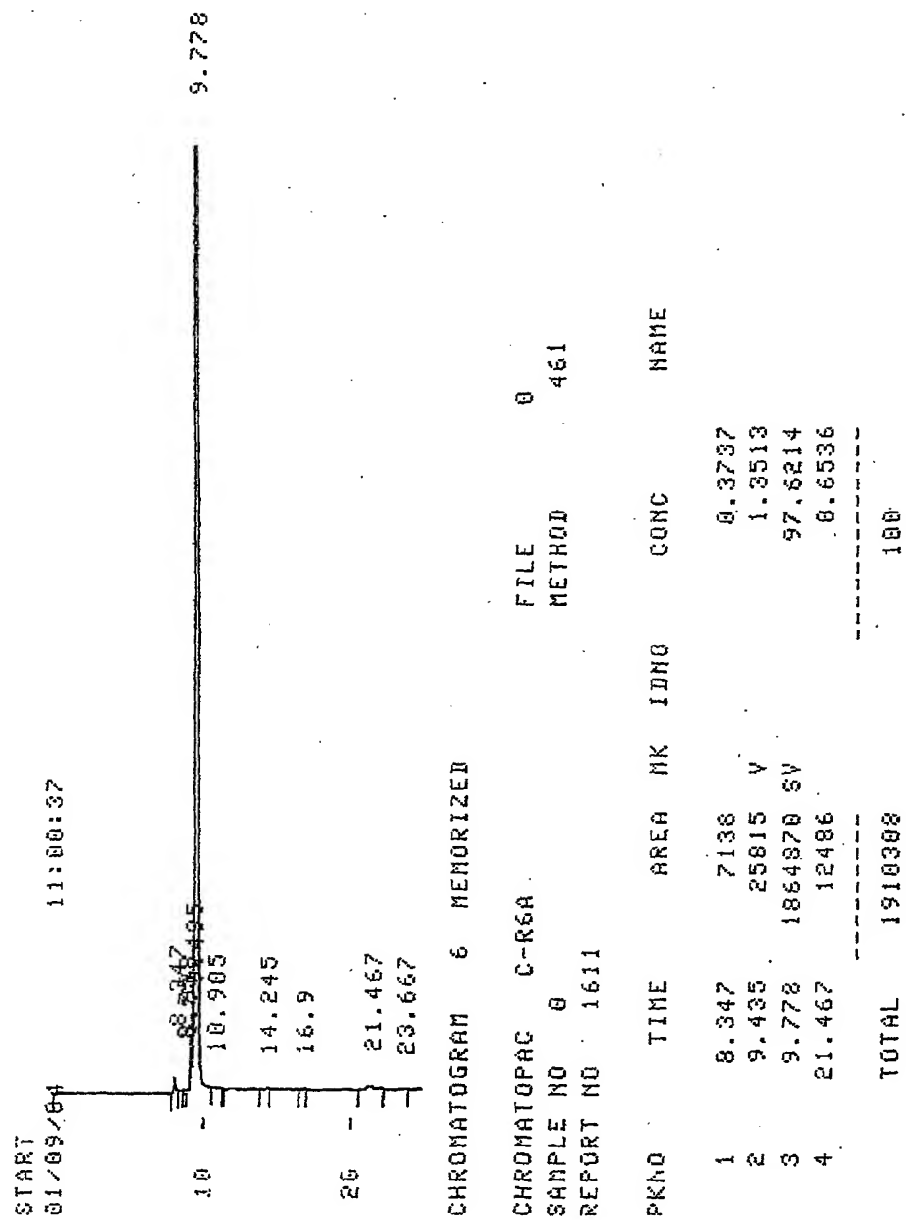
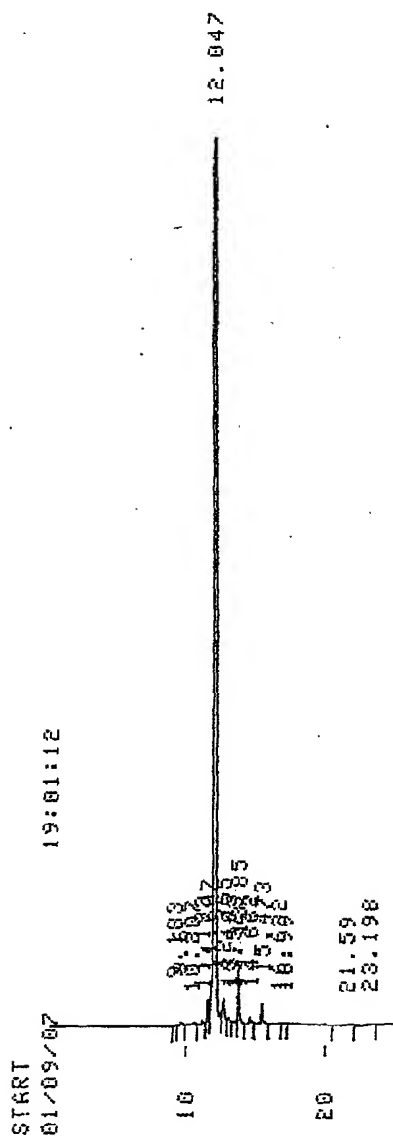
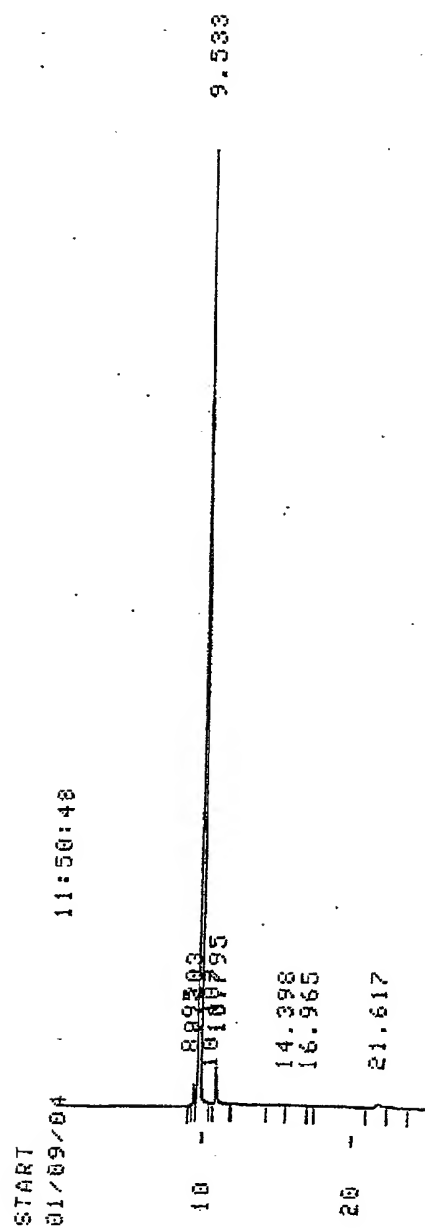


Fig. 10



PKNO	TIME	AREA	MK	IDNO	CONC	NAME
1	10.207	5378	V		0.2243	
2	11.193	5257	V		0.2193	
3	11.607	13768	V		0.5744	
4	12.047	2310076	SV		96.3713	
5	12.755	14923	T		0.6225	
6	13.85	32645	T		1.3619	
7	15.470	15012	TV		0.6263	
TOTAL					2397058	100

Fig. 11



CHROMATOGRAM 7 MEMORIZED

CHROMATOPAC C-R6A

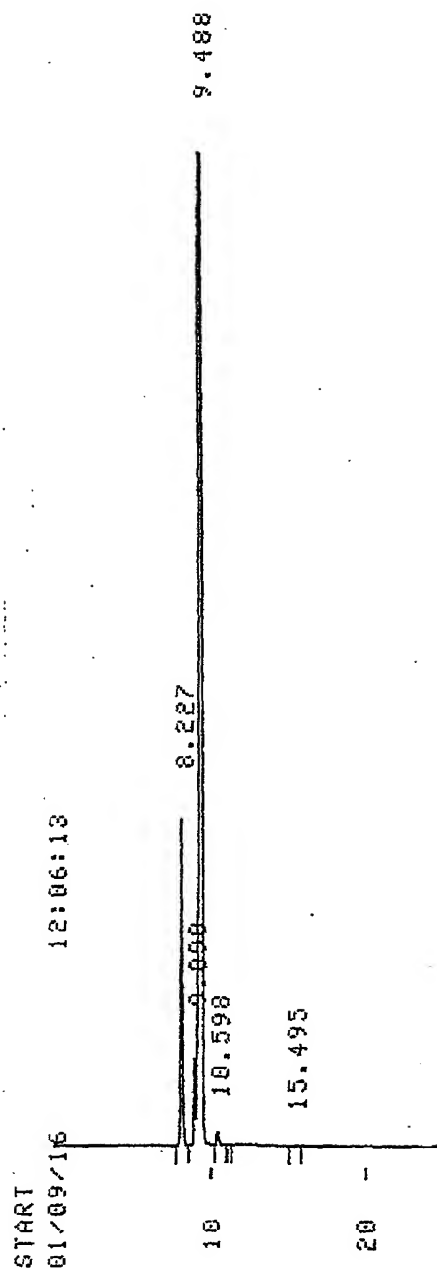
SAMPLE NO 0

REPORT NO 1612

FILE 0
METHOD 461

PKNO	TIME	AREA	AK	IDNO	CONC	NAME
1	9.203	9544	V		1.0628	
2	9.533	858444	SV		95.594	
3	10.795	21956	T		2.445	
4	21.617	8066			0.8982	
TOTAL		898010			100	

Fig. 12



CHROMATOPAC C-R6A				FILE	0	461
SAMPLE NO 0				METHOD		
REPORT NO 1652						
PKNO	TIME	AREA	MK	IDNO	CONC	NAME
1	8.227	175028			10.5415	
2	9.098	44749	V		2.6951	
3	9.488	1429985	SV		86.125	
4	10.598	10599	T		0.6383	
TOTAL					100	

Fig. 13

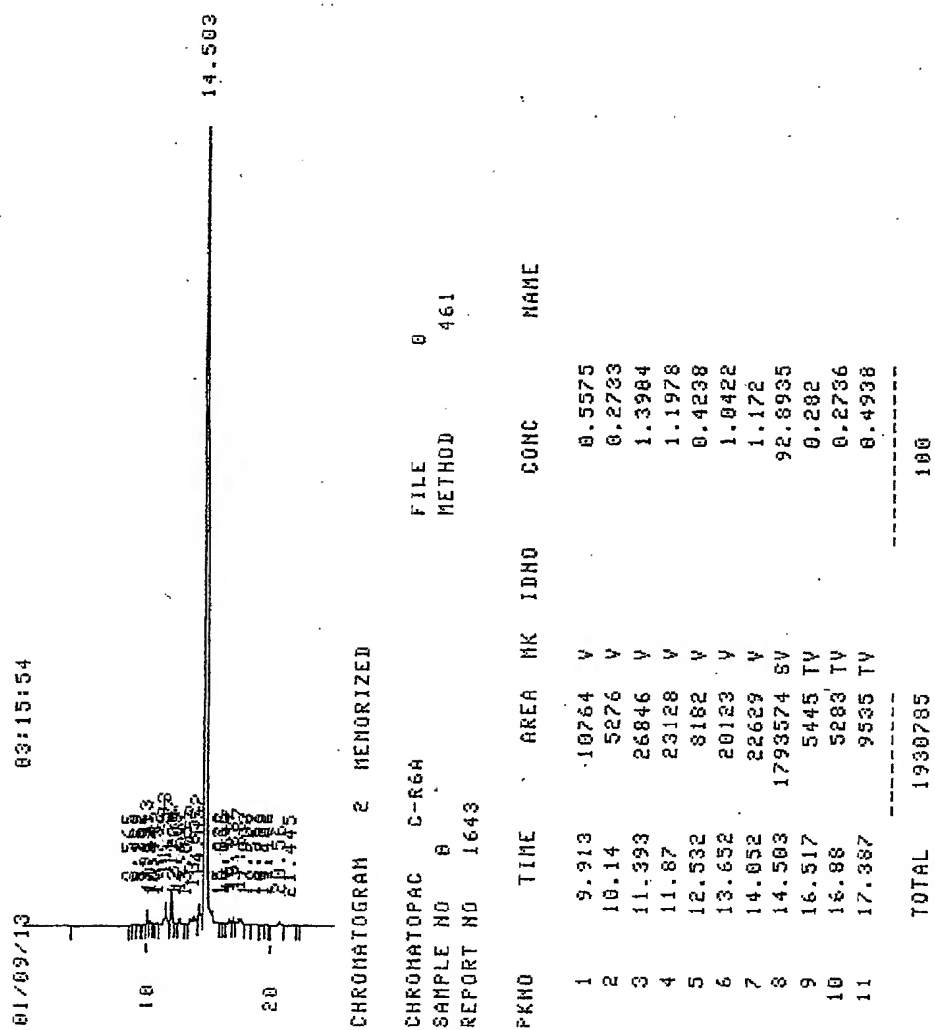
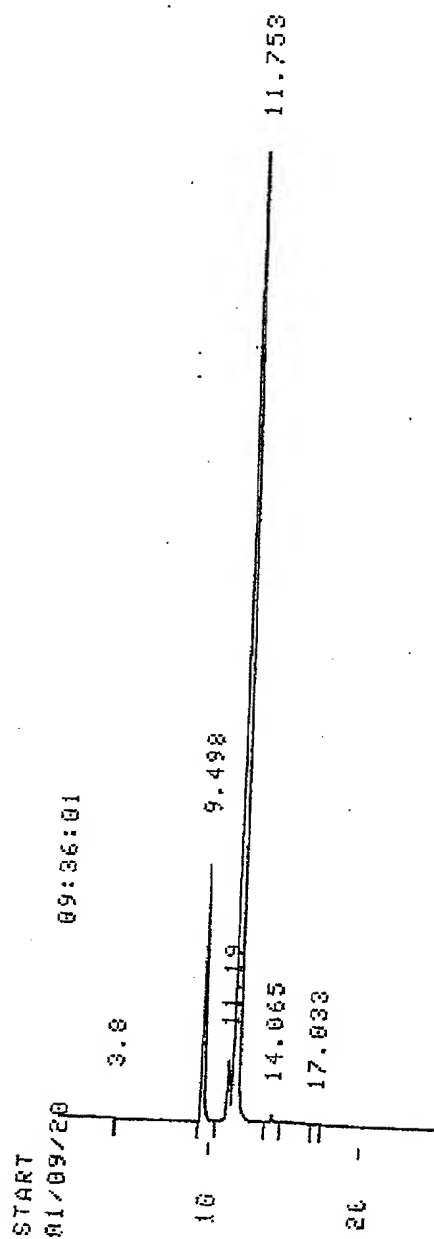


Fig. 14



CHROMATOPAC C-R6A				FILE	0	461
SAMPLE NO 0				METHOD		
REPORT NO 1662						
PKNO	TIME	AREA	MK	IDNO	CONC	NAME
1	9.498	172703			11.3246	
2	11.19	57571	V		3.775	
3	11.753	1286600	SV		84.3657	
4	14.065	8154	T		0.5347	
TOTAL		1525027			100	

Fig. 15

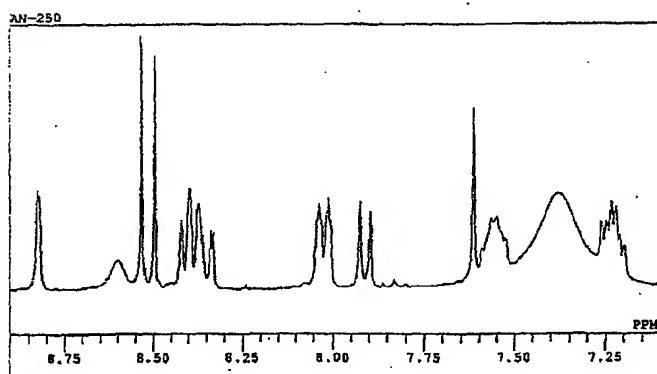
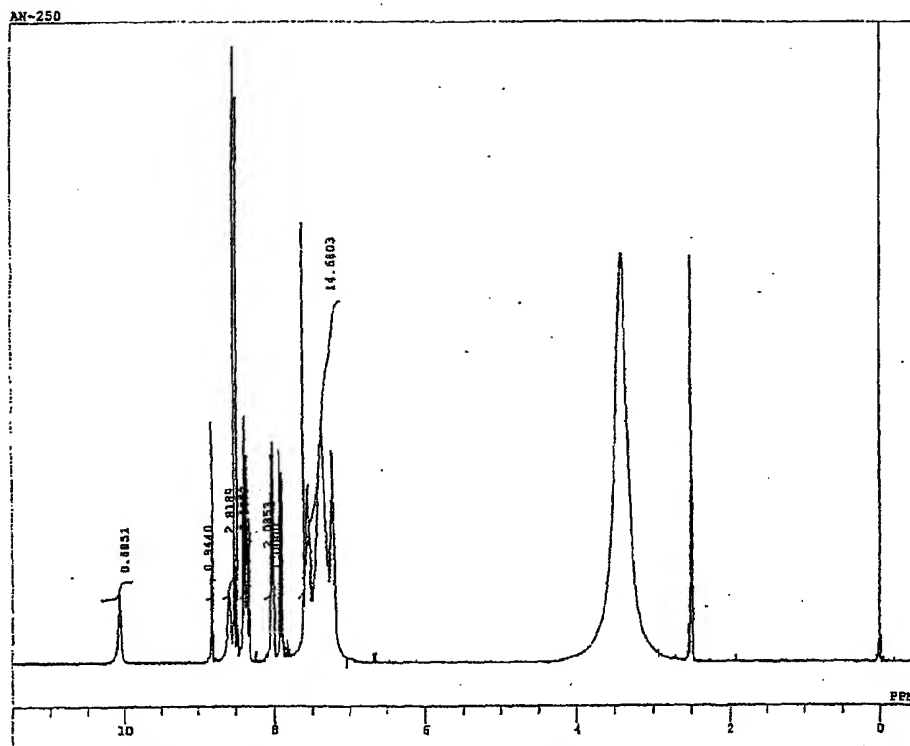


Fig. 16

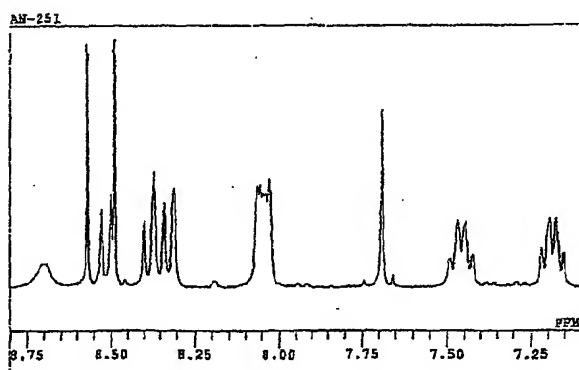
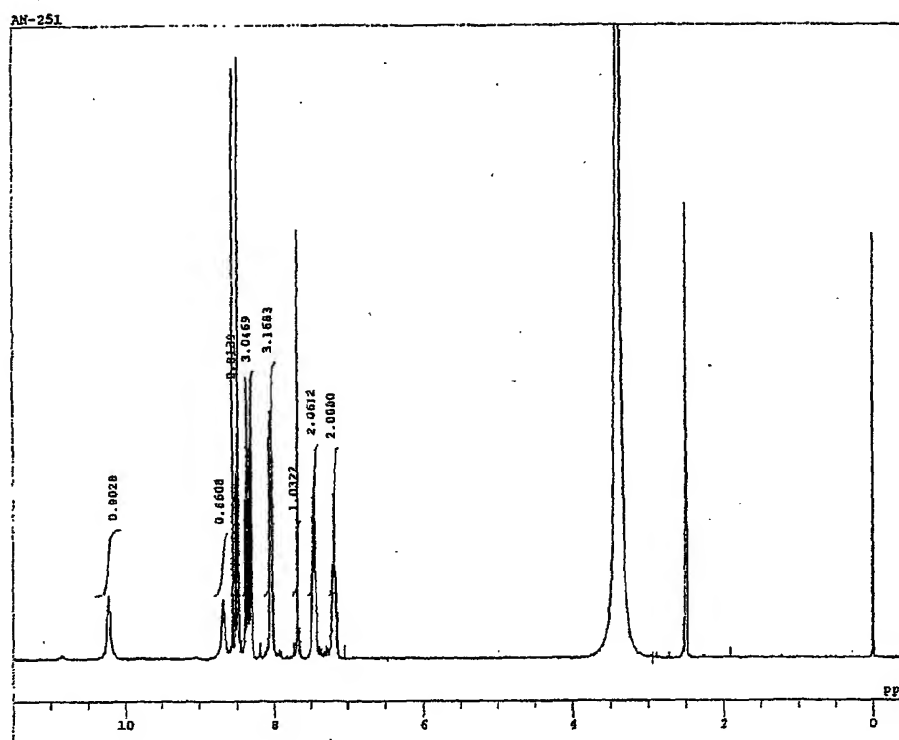


Fig. 17

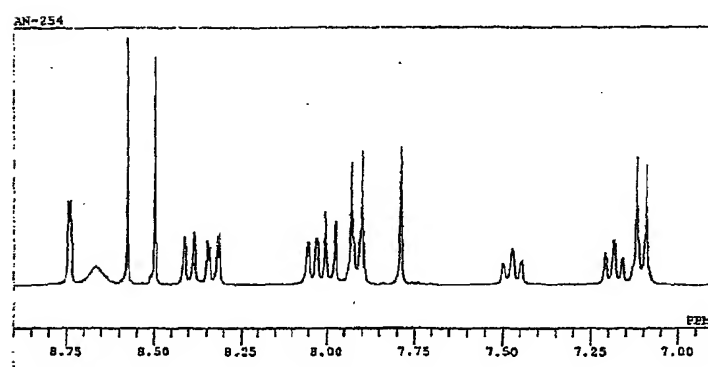
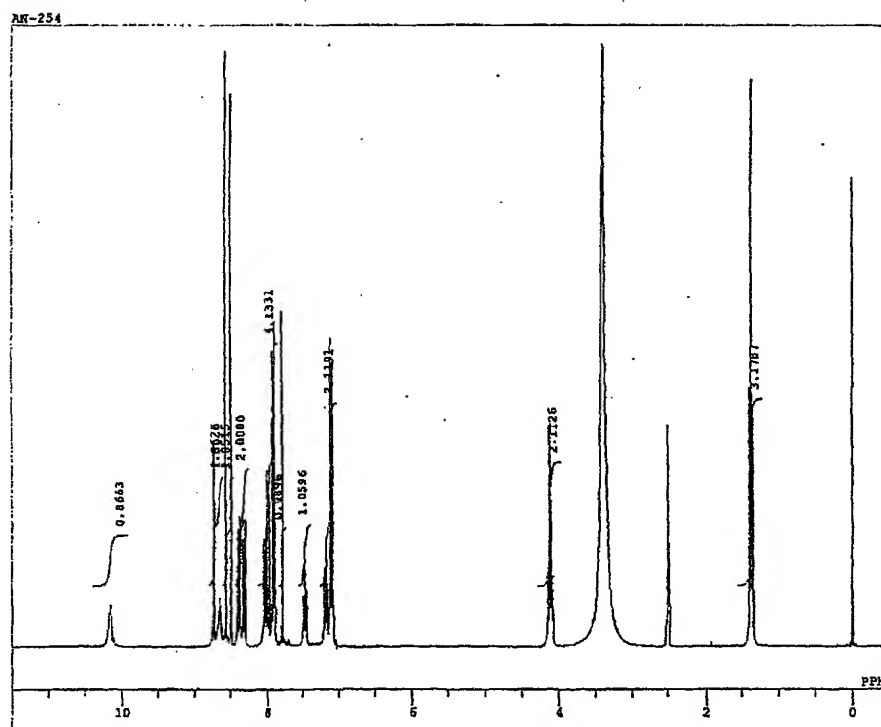


Fig. 18

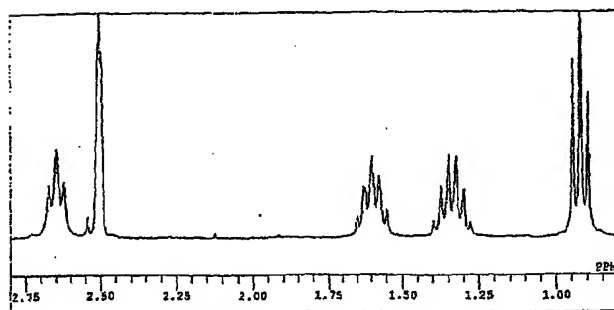
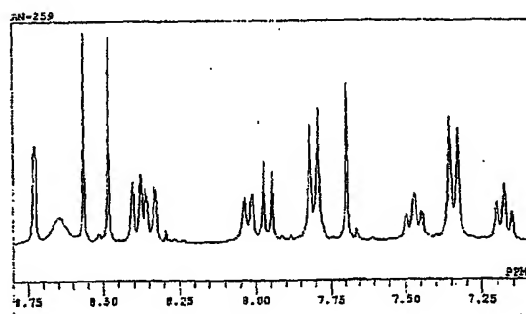
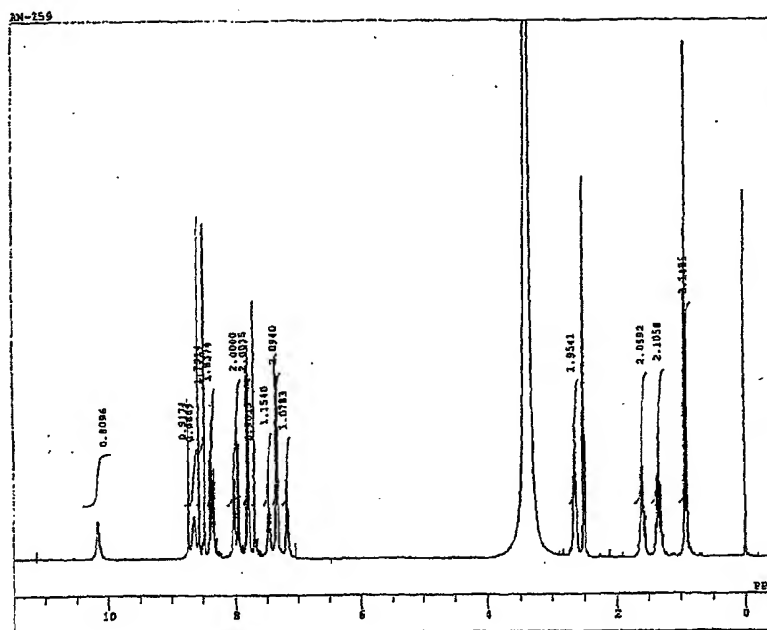


Fig. 19

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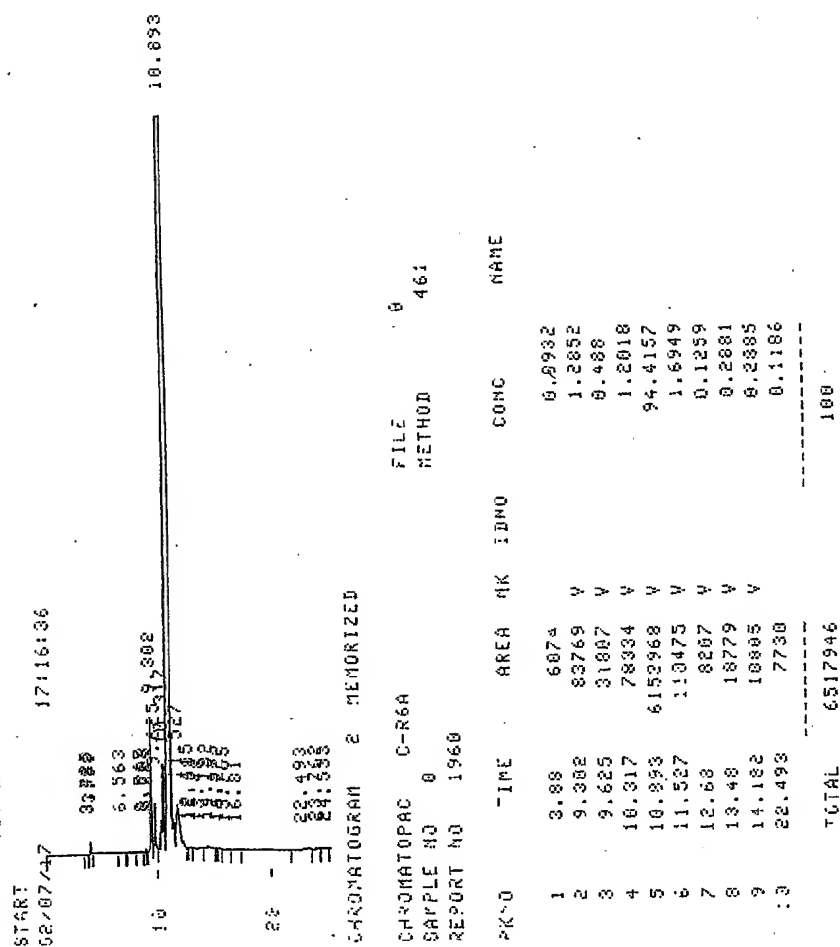
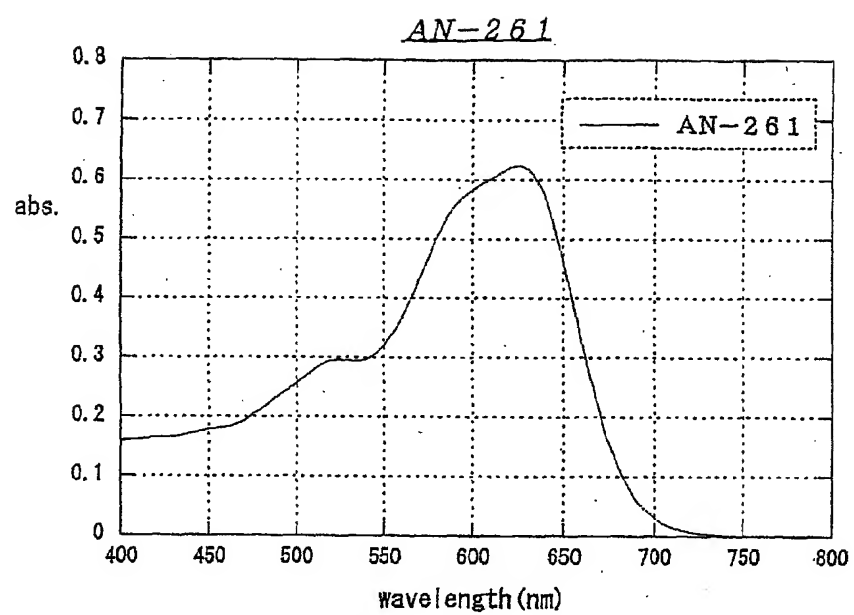


Fig. 20





European Patent
Office

EUROPEAN SEARCH REPORT

Application Number
EP 02 02 4409

DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int.Cl.7)
A	EP 0 539 178 A (LEXMARK INT INC) 28 April 1993 (1993-04-28) * claims 1,6-16; examples 5-7 *	1-7	C09B33/22 //C09D11/02
D	& JP 05 262998 A 12 October 1993 (1993-10-12) ----		
A	EP 0 645 435 A (ORIENT CHEMICAL IND) 29 March 1995 (1995-03-29) * claims; examples *	1-7	
A	EP 0 556 741 A (ORIENT CHEMICAL IND) 25 August 1993 (1993-08-25) * claims; examples * -----	1-7	
			TECHNICAL FIELDS SEARCHED (Int.Cl.7)
			C09B
The present search report has been drawn up for all claims			
Place of search THE HAGUE		Date of completion of the search 24 January 2003	Examiner Ginoux, C
CATEGORY OF CITED DOCUMENTS X : particularly relevant if taken alone Y : particularly relevant if combined with another document of the same category A : technological background O : non-written disclosure P : intermediate document		T : theory or principle underlying the invention E : earlier patent document, but published on, or after the filing date D : document cited in the application L : document cited for other reasons & : member of the same patent family, corresponding document	

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EP 02 02 4409

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24-01-2003

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
EP 0539178 A	28-04-1993	DE 69224031 D1	19-02-1998
		DE 69224031 T2	02-07-1998
		EP 0539178 A2	28-04-1993
		JP 3162834 B2	08-05-2001
		JP 5262998 A	12-10-1993
		US 5198022 A	30-03-1993
EP 0645435 A	29-03-1995	JP 3165764 B2	14-05-2001
		JP 7097529 A	11-04-1995
		DE 69417327 D1	29-04-1999
		DE 69417327 T2	04-11-1999
		EP 0645435 A1	29-03-1995
		US 5486601 A	23-01-1996
EP 0556741 A	25-08-1993	JP 2983742 B2	29-11-1999
		JP 5222299 A	31-08-1993
		DE 69306275 D1	16-01-1997
		DE 69306275 T2	03-07-1997
		EP 0556741 A1	25-08-1993
		US 5340929 A	23-08-1994

EPC FORM P0459

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